

Acid, Base, or Neutral?

Focus Questions:

1. How can you determine whether a substance is an acid, base, or neutral?
2. How can you determine the _____ of an acid or base?

Evidence: (from Interactions in Physical Science textbook)

1. Properties and Names of Acids, Bases, and Neutrals.

Directions: Read pages 421-426 to fill out Table 1 and answer the following questions.

Table 1: Some Properties and Names of Acids, Bases, and Neutral Substances

	Some common properties	Examples: List at least 8 examples each
Acids	<ol style="list-style-type: none">1.2.3.4.	<ol style="list-style-type: none">1.2.3.4.5.6.7.8.
Bases	<ol style="list-style-type: none">1.2.3.4.	<ol style="list-style-type: none">1.2.3.4.5.6.7.8.
Neutrals	<ol style="list-style-type: none">1. Do _____ react with indicators2. Have a wide variety of different properties.	<ol style="list-style-type: none">1.2.3.4.5.6.7.8.

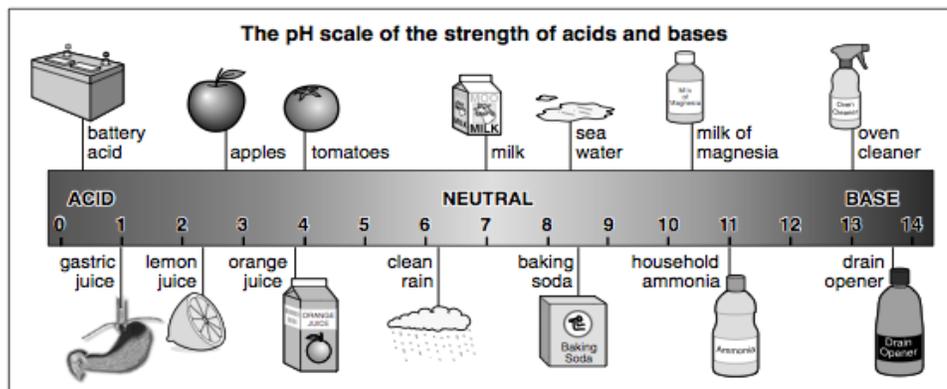
2. Cornell Notes: Pages 424 (Complete sentences for full credit)

1. What happens when an acid is mixed with a base?	
2. What is neutralization and what does neutralization always result in?	

3. What is a SALT ?	
4. What is the pH scale ?	
5. How can you tell if a substance is an acid? A base? A neutral? (explain the pH scale)	
6. What is a " <u>corrosive</u> " acid and what can it do to substances?	
7. What is a " <u>corrosive</u> " base and what can it do to substances?	

3. "Make Sense of the Strength of Acids and Bases"

Directions: Read page 424: "The pH Scale" and use the pH scale (below) to answer the following questions:



1. Which two acids shown on the pH scale are corrosive? **Explain** your answer.

2. Which two bases shown on the pH scale are corrosive? **Explain** your answer.

3. A brand of hair remover has a pH number of 12.3. Is this brand of hair remover acidic, neutral, or basic? Is it considered corrosive? **Explain** your answer.

4. The saliva in your mouth has a pH number of 7.1. Is your saliva acidic, neutral, or basic?

5. Which acid is stronger, the acid in apples or the acid in orange juice? **Explain** your answer.